**Writing chemical equations**

**For each of these reactions complete the table by:**

* 1. **writing a word equation**
  2. **writing formulas for the reactants and products (hint: check the formulas sheet)**
  3. **writing a BALANCED formula equation.**

1. Solid carbon reacts with water vapour to produce carbon monoxide gas and hydrogen gas.

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| **Word equation** |  |  |
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| **Formulas** | **Reactants** | **Products** |
|  |  |  |
| **Chemical equation** |  |  |
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1. Zinc metal reacts with copper sulfate solution to produce copper metal and zinc sulfate solution.

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| **Word equation** |  |  |
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| **Formulas** | **Reactants** | **Products** |
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| **Chemical equation** |  |  |
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1. Sodium chloride solution reacts with silver nitrate solution, to produce solid silver chloride and sodium nitrate solution

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| **Word equation** |  |  |
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| **Formulas** | **Reactants** | **Products** |
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| **Chemical equation** |  |  |
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1. Zinc metal reacts with dilute sulfuric acid to produce zinc sulfate solution and hydrogen gas.

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| **Word equation** |  |  |
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| **Formulas** | **Reactants** | **Products** |
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| **Chemical equation** |  |  |
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1. Solid copper (II) oxide reacts with hydrogen gas to produce copper metal and water vapour.

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| **Word equation** |  |  |
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| **Formulas** | **Reactants** | **Products** |
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| **Chemical equation** |  |  |
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1. Solid magnesium carbonate, when heated, yields solid magnesium oxide and carbon dioxide gas.

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| **Word equation** |  |  |
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| **Formulas** | **Reactants** | **Products** |
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| **Chemical equation** |  |  |
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1. Calcium carbonate solution reacts with dilute sulfuric acid to produce sodium chloride, carbon dioxide gas and liquid water.

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| **Word equation** |  |  |
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| **Formulas** | **Reactants** | **Products** |
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| **Chemical equation** |  |  |
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1. Potassium hydroxide solution reacts with nitric acid solution to produce potassium nitrate solution and liquid water.

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| **Word equation** |  |  |
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| **Formulas** | **Reactants** | **Products** |
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| **Chemical equation** |  |  |
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**9.** Ammonia gas reacts with hydrogen chloride gas to produce ammonium chloride solution.

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| **Word equation** |  |  |
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| **Formulas** | **Reactants** | **Products** |
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| **Chemical equation** |  |  |
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